Fill in the blanks in the table by **calculating** the missing values:

Mass of solute	Volume of liquid	Concentration mol / dm^3	Concentration g/dm^3
NaOH = 23+17 = 40 50 g of NaOH $\frac{50}{40} = 1.25$	4 dm^3	NaOH = 40 $\frac{50}{40} = 1.25$ $\frac{1.25}{4} = 0.3125$	12.5 g / dm^3
12 g of $MgCl_2$ 95 $\frac{12}{95} = 0.1263...$	0.084 dm^3 8 dm^3	1.5 mol / dm^3	142.5 g / dm^3
KCl = 74.5 8 g of KCl $\frac{8}{74.5} = 0.10738...$	0.13 dm^3 10 dm^3	0.8 mol / dm^3	59.6 g / dm^3
$Mg(OH)_2$ = 58 98 g of $Mg(OH)_2$ $\frac{98}{58} = 1.689...$	5 dm^3	0.34 mol / dm^3	17.6 g / dm^3