

$$I = \frac{1}{2} \cdot \sum_i Z_i^2 \cdot c_i$$

For compound B_nA_m coefficient t equals to:

$$t = \frac{1}{2} \cdot (n \cdot Z_B^2 + m \cdot Z_A^2)$$

$$I = t \cdot c$$

Electrolyte type*	Compound	Ionic strength as a function of concentration ($t \cdot c$)
1-1	KCl or NaH_2PO_4	$1 \cdot c$
1-2	$CaCl_2$ or Na_2SO_4	$3 \cdot c$
2-2	$MgSO_4$	$4 \cdot c$
1-3	Na_3PO_4 or $Fe(NO_3)_3$	$6 \cdot c$
1-2-3	$KAl(SO_4)_2$	$9 \cdot c$
2-3	$Cr_2(SO_4)_3$	$15 \cdot c$